

TYPES OF ORGANIC COMPOUNDS

HYDROCARBON	HALO-CARBON	HALO-CARBON	HALO-CARBON	NITROGEN ADDED	SULFUR ADDED
ALKANE 	ALCOHOL $-OH$	ALCOHOL $-OH$	CARBOXYLIC ACID 	AMINE $-NH_2$	thiol: SH
- ethane: C_2H_6	- methanol: CH_3OH (- methyl alcohol)	- ethanol: CH_3CH_2OH	- ethanoic acid: CH_3COOH (acetic acid)	- methyl amine: CH_3NH_2	thioether: $R-S-R'$
- methyl: CH_3	- chloro: CH_3Cl	- chloro: CH_3Cl	- acetate ion: CH_3COO^-	- ethylamine: CH_3NH_2	disulfide: $R-S-S-R'$
ALKENE 	- diol: CH_2OH	- diol: CH_2OH	- boronic acid: $CH_3COOB(OH)_2$	- thiolate: $CH_3NH_2^-$	thiolide: $R-SO-R'$
- ethene: C_2H_4	- glycol: CH_2OH_2	- borontrifluoride: CH_3COBF_3	Dimethyl ester: $CH_3COOC_2H_5$	- sulfide: CH_3NH_2H	disulfide: $R-SO_2-R'$
- diene: $CH_2=CH-C=CH$	- ether: CH_2OCH_3	- 2-propanone: CH_3COCH_3 (dimethyl ketone, acetone)	PROX: $CH_3COOC_2H_5$	- nitrile: $CH_3C≡N$	nitro: $R-NO_2$
ALKYNE 	- ethoxyethane: $CH_3OCH_2CH_3$ (- ethyl ether)	- diketone: $R-C(=O)R-C(=O)R$	ESTER 	- carbamide: CH_3CONH_2	nitroso: $R-NO$
- ethyne: C_2H_2	- epoxy: $CH_2OCH=C=CH_2$		Other derivatives	- amide: CH_3CONH_2	halogen: $-X$
AROMATIC 	- cyclic ether		- Pernitro: CH_3COONO_2	- acid amide: $CH_3CO-O-CH_2CH_3$	chloroform: $CHCl_3$
- benzene: C_6H_6	PEROXIDE: $R-O-O-R'$		- Acid anhydride: $CH_3CO-O-CH_2CO$	- nitrile oxide: $CH_3C≡N-O$	acyl: CH_3CO-
- aniline: $C_6H_5NH_2$ ($Ar-NH_2$)					acyl fluoride: CH_3CO-F
					acyl sulfide: CH_3CO-SH

FORMULAS AND ISOMERS

Molecular formulas: Elementary symbols with subscripts denote the composition of a compound
Empirical formulas: Subscripts denote the relative elemental composition

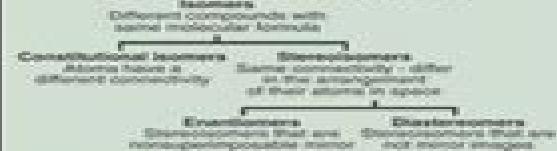
Graphical depiction:

- **Dash formula:** Diagrams all atoms, bonds as dashes.
- **Bond line formula:** Shows H, shows carbon skeletons as lines, other atoms explicit

• **Newman Projections:** 2-D depiction

• **3-Dimensional:** Models of molecule's actual structure

SUBDIVISIONS OF ISOMERS



constitutional isomers: different bonding connectivity (e.g. rings, bonds, branching, substituted positions)
tautomeric isomers: interconverted structural isomers (e.g. keto-enol for ketones)

chiral: not identical with mirror image

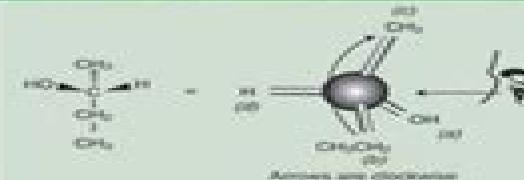
achiral: has plane of symmetry (superimposable on mirror image)

enantiomers: a pair of diastereomers which differ only in the configuration of one atom

More than 1 chiral center:

- in chiral systems, $\leq 2^n$ stereoisomers
- more than chiral centers, 4 isomers: 3 stereoisomers, 1 achiral (enantiomer)

FORMULAS AND ISOMERS



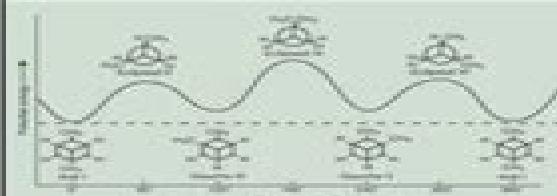
R/S notation: The four different atoms or groups attached to a central atom are ranked a,b,c,d by molecular mass. The lowest (d) is directed away from the viewer and the sequence of a-b-c produces clockwise (R) or counter-clockwise (S) configuration.

- chiral (optically active): + or - rotation of plane polarized light, R/S: Opposite effects
- meso: 50:50 mixture of enantiomers (no net optical activity)
- anomers: None R/S and +/-. In the compound name, example: $2R,3S$ bromochloroethanol.

Fischer-projection: Diagram depicts chiral 3-D structure
• molecular conformations: molecule exhibits structural variations due to free rotation about C-C single bond.



Newman-diagram: Depict rotation about a C-C bond; aligned (high energy), anti (low energy), gauche (intermediate energy).



COMMON TERMS

- **aliphatic:** non-aromatic
- **aromatic:** benzene ring
- **heterocyclic:** non-carbon atom in the ring structure
- **hydrocarbon:** compound of H and C
- **paraffin:** alkane
- **olefin:** alkene
- **saturated:** contains ≤ 2 SP² (all C-C single bonds)
- **unsaturated:** at least one C-C multiple bond

NOMENCLATURE

IUPAC - standard guidelines for naming compounds.
Nomenclature Strategy: Find longest carbon chain, identify and note location of functional groups and substituents by chain position number.

Classes of compounds: are defined by the functional group. There are many common names and functional group names. Multiple names are possible.

CARBON CHAIN PREFIXES

CARBON CHAIN	PREFIX	PREFIX
1	meth-	meth-
2	eth-	eth-
3	prop-	prop-
4	but-	but-
5	pent-	pent-
6	hex-	hex-
7	hept-	hept-
8	oct-	oct-
9	non-	non-
10	dec-	dec-

cyclic: ring structure; example: cyclohexane: 6-carbon ring molecule

iso: two methyl groups on the same carbon

tertiary: carbon with three substituents

sec: vicinally: two substituents on adjacent carbons

tertiary: two substituents on the same carbon

aromatic isomers: cis or trans

hetero: substitution positions:

ortho (1,2), meta (1,3), para (1,4)

Carbon Position



Carbon atoms & H-atoms

1: 6 carbons, 12 H-atoms

2: 5 carbons, 10 H-atoms

3: 4 carbons, 8 H-atoms

4: 3 carbons, 6 H-atoms

5: 2 carbons, 4 H-atoms

6: 1 carbon, 2 H-atoms

Organic Chemistry Study Cards 3

Zhe Xu

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