

TYPES OF ORGANIC COMPOUNDS

HYDROCARBON	O-ADDED	C=O ADDED	COO ADDED	NITROGEN ADDED	SULFUR ADDED
ALKANE C_nH_{2n+2} <ul style="list-style-type: none"> ethane: C_2H_6 methyl (Me): $-CH_3$ ethyl (Et): $-C_2H_5$ ALKENE C_nH_{2n} <ul style="list-style-type: none"> ethene: C_2H_4 diene: two C=C triene: three C=C ALKYNE C_nH_{2n-2} <ul style="list-style-type: none"> ethyne: C_2H_2 AROMATIC C_6H_6 <ul style="list-style-type: none"> benzene: C_6H_6 aromatic: C_6H_5 (Ar-) 	ALCOHOL $C_nH_{2n+2}O$ <ul style="list-style-type: none"> methanol: $Me-OH$ (methyl alcohol) phenol: $Ar-OH$ diol/glycol: (2-OH) glycerol: (3-OH) ETHER $C_nH_{2n+2}O$ <ul style="list-style-type: none"> ethoxyethane: $Et-O-Et$ (diethyl ether) EPOXY <ul style="list-style-type: none"> cyclic ether PEROXIDE $R-O-O-R'$	ALDEHYDE <ul style="list-style-type: none"> methanal: H_2CO (formaldehyde) hexaldehyde: $Ar-CHO$ KETONE <ul style="list-style-type: none"> 2-propanone: $Me-CO-Me$ (dimethyl ketone, acetone) diketone: $R-CO-R'-CO-R''$ 	CARBOXYLIC ACID <ul style="list-style-type: none"> ethanoic acid: $Me-COOH$ (acetic acid) acetic acid: $Me-COOH$ benzoic acid: $Ar-COOH$ Dicarboxylic acid $HOOC-R-COOH$ ESTER <ul style="list-style-type: none"> ethyl acetate: $Me-CO-OEt$ Other derivatives: <ul style="list-style-type: none"> Formic acid: $H-CO-OH$ Acid anhydride: $RCO-O-CO-R'$ 	AMINE <ul style="list-style-type: none"> methyl amine: H_3C-NH_2 phenyl amine: $Ar-NH_2$ (aniline) $R-NH_2$ (1°), $RR'NH$ (2°), $RR'R''N$ (3°) NITRO $R-NO_2$ DIAZO $R-N=N$ NITRILE $R-CN$ <ul style="list-style-type: none"> methane nitrile: $Me-CN$ AMIDE <ul style="list-style-type: none"> acetamide: $Me-CO-NH_2$ 	<ul style="list-style-type: none"> thiol: $R-SH$ thioether: $R-S-R'$ disulfide: $R-S-S-R'$ thiol ester: $R-CO-SR'$ sulfide: $R-SO-R'$ sulfone: $R-SO_2-R'$ sulfonic acid: $R-SO_3H$ HALOGEN ADDED <ul style="list-style-type: none"> haloalkane: $R-X$ (Me-Cl chloromethane) halobenzene: $Ar-X$ (chlorobenzene: $Ar-Cl$) acyl halide: $R-CO-X$ aryl halide: $Ar-X$

FORMULAS AND ISOMERS

Molecular formula: Elemental symbols with subscripts denote the composition of a compound.
Empirical formula: Subscripts denote the relative elemental composition.

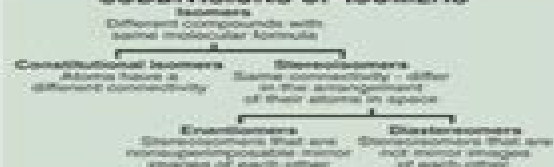
Graphical depiction:

- Dash formula:** Diagram all atoms, bonds as dashes.
- Bond line formula:** Hide H, show carbon skeleton as lines, other atoms explicit.



- Newman Projection:** 2-d depiction
- 3-dimensional:** Wedges of sawhorse denote structure

SUBDIVISIONS OF ISOMERS



constitutional isomers: different bonding connectivity (ex. rings, bonds, branching, substituent positions)
tautomers: easily interconverted structural isomers (ex. keto-enol for ketones)

chiral: not identical with mirror image
achiral: has plane of symmetry (superimposable on mirror-image)
enantiomers: a pair of diastereomers which differ only in the configuration of one atom

More than 1 chiral center:

- n chiral centers, 2^n stereoisomers
- meso:** two chiral centers, 4 isomers: 3 stereoisomers, 1 achiral (mirror-plane)

FORMULAS AND ISOMERS



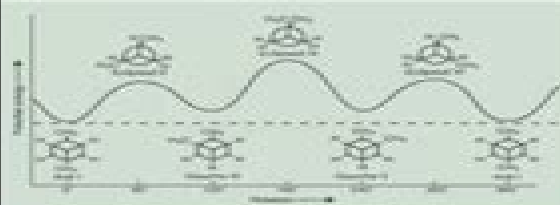
R/S notation: The four different atoms or groups attached to a central atom are ranked a,b,c,d, by molar mass. The lowest (d) is directed away from the viewer and the sequence of a-b-c produces clockwise (R) or counter-clockwise (S) configuration.

- chiral (optically active):** + or - rotation of plane polarized light, R/S: Opposite effects
- racemic:** 50:50 mixture of stereoisomers (no net optical activity)
- nomenclature:** Note R/S and +/- in the compound name, example: R (-)- bromochloromethanol.

Fischer projection: Diagram depicts chiral 3-D structure
molecular conformations: molecule exhibits structural variation due to free rotation about C-C single bond.



Newman diagram: Depict rotation about a C-C bond; eclipsed (high energy), anti (low energy), gauche (intermediate energy).



COMMON TERMS

aliphatic: non-aromatic
aromatic: benzene ring
heterocyclic: non-carbon atom in the ring structure
hydrocarbon: compound of H and C
paraffin: alkane
olefin: alkene
saturated: maximum # of H's (all C-C single bonds)
unsaturated: at least one C-C multiple bond

NOMENCLATURE

IUPAC - standard guidelines for naming compounds.
Nomenclature Strategy - find longest carbon chain, identify and note location of functional groups and substituents by chain position number.

Classes of compounds are defined by the functional group. There are many common names and functional group names. Multiple names are possible.

CARBON CHAIN PREFIXES

Ends	Ends	Ends
1	methyl	ethyl
2	ethyl	propyl
3	propyl	butyl
4	butyl	pentyl
5	pentyl	hexyl
6	hexyl	heptyl
7	heptyl	octyl
8	octyl	nonyl
9	nonyl	decyl
10	decyl	

cyclo- ring structure; example: cyclopentane 3-carbon ring molecule

iso- two methyl groups on the

terminal of a chain

n- normal straight chain

i- tertiary alkyl group

vic (vicinally) two substituents

on adjacent carbons

gem (geminally) two substituents

on the same carbon

alkene isomers: cis or trans

benzene substitution positions:

ortho (1,2), meta (1,3), para (1,4)

Carbon Position



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