

P Block Elements

valence electron configuration is $ns^2 np^{1-6}$

The periodic table is shown with the P Block elements highlighted in pink boxes. The pink boxes are located in the following positions:

- Row 2: Elements B, C, N, O, F, Ne
- Row 3: Elements Al, Si, P, S, Cl, Ar
- Row 4: Elements Ga, Ge, As, Se, Br, Kr
- Row 5: Elements In, Sn, Sb, Te, I, Xe
- Row 6: Elements Tl, Pb, Bi, Po, At, Rn
- Row 7: Elements Nh, Fl, Mc, Lv, Ts, Og
- Row 8: Elements La, Ce, Pr, Nd, Pm, Sm, Eu, Gd, Tb, Dy, Ho, Er, Tm, Yb, Lu
- Row 9: Elements Ac, Th, Pa, U, Np, Pu, Am, Cm, Bk, Cf, Es, Fm, Md, No, Lr

The table also includes the following information:

- Group 18:** He (atomic number 2, atomic mass 4.003)
- Group 17:** F (atomic number 9, atomic mass 18.015), Cl (atomic number 17, atomic mass 35.453), Br (atomic number 35, atomic mass 79.904), I (atomic number 53, atomic mass 126.904), Xe (atomic number 54, atomic mass 131.296)
- Group 16:** O (atomic number 8, atomic mass 15.999), S (atomic number 16, atomic mass 32.065), Se (atomic number 34, atomic mass 78.976), Te (atomic number 52, atomic mass 127.905)
- Group 15:** N (atomic number 7, atomic mass 14.007), P (atomic number 15, atomic mass 30.973), As (atomic number 33, atomic mass 74.924), Sb (atomic number 51, atomic mass 121.770)
- Group 14:** C (atomic number 6, atomic mass 12.011), Si (atomic number 14, atomic mass 28.085), Ge (atomic number 32, atomic mass 69.724), Sn (atomic number 50, atomic mass 113.785)
- Group 13:** B (atomic number 5, atomic mass 10.811), Al (atomic number 13, atomic mass 26.982), Ga (atomic number 31, atomic mass 69.724), In (atomic number 49, atomic mass 114.818)
- Group 12:** Ca (atomic number 20, atomic mass 40.078), Sr (atomic number 38, atomic mass 87.675), Ba (atomic number 56, atomic mass 137.332)
- Group 11:** K (atomic number 19, atomic mass 39.098), Rb (atomic number 37, atomic mass 85.468), Cs (atomic number 55, atomic mass 132.910)
- Group 10:** Ca (atomic number 20, atomic mass 40.078), Sr (atomic number 38, atomic mass 87.675), Ba (atomic number 56, atomic mass 137.332)
- Group 9:** Ti (atomic number 22, atomic mass 47.867), V (atomic number 23, atomic mass 50.942), Cr (atomic number 24, atomic mass 51.996), Mn (atomic number 25, atomic mass 54.938), Fe (atomic number 26, atomic mass 55.845), Co (atomic number 27, atomic mass 58.933), Ni (atomic number 28, atomic mass 58.693), Cu (atomic number 29, atomic mass 63.546), Zn (atomic number 30, atomic mass 65.401)
- Group 8:** Cr (atomic number 24, atomic mass 51.996), Mn (atomic number 25, atomic mass 54.938), Fe (atomic number 26, atomic mass 55.845), Co (atomic number 27, atomic mass 58.933), Ni (atomic number 28, atomic mass 58.693), Cu (atomic number 29, atomic mass 63.546), Zn (atomic number 30, atomic mass 65.401)
- Group 7:** V (atomic number 23, atomic mass 50.942), Cr (atomic number 24, atomic mass 51.996), Mn (atomic number 25, atomic mass 54.938), Fe (atomic number 26, atomic mass 55.845), Co (atomic number 27, atomic mass 58.933), Ni (atomic number 28, atomic mass 58.693), Cu (atomic number 29, atomic mass 63.546), Zn (atomic number 30, atomic mass 65.401)
- Group 6:** Ti (atomic number 22, atomic mass 47.867), V (atomic number 23, atomic mass 50.942), Cr (atomic number 24, atomic mass 51.996), Mn (atomic number 25, atomic mass 54.938), Fe (atomic number 26, atomic mass 55.845), Co (atomic number 27, atomic mass 58.933), Ni (atomic number 28, atomic mass 58.693), Cu (atomic number 29, atomic mass 63.546), Zn (atomic number 30, atomic mass 65.401)
- Group 5:** Sc (atomic number 21, atomic mass 44.959), Ti (atomic number 22, atomic mass 47.867), V (atomic number 23, atomic mass 50.942), Cr (atomic number 24, atomic mass 51.996), Mn (atomic number 25, atomic mass 54.938), Fe (atomic number 26, atomic mass 55.845), Co (atomic number 27, atomic mass 58.933), Ni (atomic number 28, atomic mass 58.693), Cu (atomic number 29, atomic mass 63.546), Zn (atomic number 30, atomic mass 65.401)
- Group 4:** Ca (atomic number 20, atomic mass 40.078), Sc (atomic number 21, atomic mass 44.959), Ti (atomic number 22, atomic mass 47.867), V (atomic number 23, atomic mass 50.942), Cr (atomic number 24, atomic mass 51.996), Mn (atomic number 25, atomic mass 54.938), Fe (atomic number 26, atomic mass 55.845), Co (atomic number 27, atomic mass 58.933), Ni (atomic number 28, atomic mass 58.693), Cu (atomic number 29, atomic mass 63.546), Zn (atomic number 30, atomic mass 65.401)
- Group 3:** Li (atomic number 3, atomic mass 6.941), Be (atomic number 4, atomic mass 9.012), Na (atomic number 11, atomic mass 22.989), Mg (atomic number 12, atomic mass 24.305)
- Group 2:** Li (atomic number 3, atomic mass 6.941), Be (atomic number 4, atomic mass 9.012), Na (atomic number 11, atomic mass 22.989), Mg (atomic number 12, atomic mass 24.305)
- Group 1:** H (atomic number 1, atomic mass 1.008)
- Group 0:** He (atomic number 2, atomic mass 4.003)

P Block Elements

S.K. Jain & Shailesh K. Jain

P Block Elements:

Periodicity and the S- and P- Block Elements Nicholas C. Norman,2021 The renowned Oxford Chemistry Primers series which provides focused introductions to a range of important topics in chemistry has been refreshed and updated to suit the needs of today's students, lecturers and postgraduate researchers. The rigorous yet accessible treatment of each subject area is ideal for those wanting a primer in a given topic to prepare them for more advanced study or research. Moreover, cutting edge examples and applications throughout the texts show the relevance of the chemistry being described to current research and industry. The learning features provided, including end of chapter questions and online multiple choice questions, encourage active learning and promote understanding. Furthermore, frequent diagrams, margin notes, further reading and glossary definitions all help to enhance a student's understanding of these essential areas of chemistry. This new and updated edition of Periodicity and the s and p Block Elements provides a compelling and accessible introduction to key periodic trends found within the s and p blocks of the periodic table and includes coverage of the elements themselves as well as the compounds they form. Additional chapters focus on acidity and basicity as well as on structure. The final chapter is entirely new to the second edition and contains a critical examination of many theories, models and approaches to the study of the ideas explored in the book. Digital formats and resources. The second edition is available for students and institutions to purchase in a variety of formats and is supported by online resources. The e-book offers a mobile experience and convenient access along with functionality, tools, navigation features and links that offer extra learning support. www.oxfordtextbooks.co.uk/ebooks Online resources include multiple choice questions for students to check their understanding and for registered adopters figures and tables from the book. **Chemistry of S and P Block Elements**, 2017 **P-block elements**, 2016

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Chemistry Allan Blackman,2010 **Conceptual Chemistry Volume I For Class XI** S.K. Jain & Shailesh K. Jain,1998

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