

• STARTING A VEGETABLE GARDEN? •

HERE'S SOME TOP TIPS



LIGHT

Track the light of where you are wanting to plant. This helps you see how much shade or sunshine somewhere gets throughout the day before planting.



GROW

Grow what you love to eat and use the most in your kitchen at first. Then experiment later on with more experience. This will help you find your feet and not get overwhelmed.



SEEDS

Seeds or starts? Growing from seed is cheaper and in most cases, leads to a healthier plant. But plant starts are a great way to step into your first time vegetable garden.



PLANTING

If you are stuck for space, you can plant a lot of things in containers and pots. Just provide them with good drainage and check the depth and width the plant needs.



MAINTENANCE

A vegetable garden can go from simple to challenging really quickly. Plant for the time you can give it. Things like root vegetables and brassicas need little effort.



SOIL

Soil is the backbone of your garden. Invest in it more than anything. Great soil will give you the head start you need. Dig or no dig? Decide what's best for you and your space.

Gardening Tips Quick Start

Guillaume Favre

A red circular graphic with a gradient, appearing as a stylized arrow or a partial circle, located to the right of the author's name.

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web jun 14 2015 this list of electron configurations of elements contains all the elements in increasing order of atomic number to save room the configurations are in noble gas shorthand this means part of the electron configuration has been replaced with the element symbol of the noble gas symbol

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web elements in the second row of the periodic table place their electrons in the $2n$ shell as well as the $1n$ shell for instance lithium li has three electrons two fill the $1s$ orbital and the third is placed in the $2s$ orbital giving an electron configuration of $1s^2 2s^1$

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web electron configurations and trends answer these discussion questions given the following part of an electron configuration $3s^2$ explain which part refers to the number of electrons the energy level and the sublevel what is the maximum number of electrons that can occupy a single orbital state and explain hund's rule

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web one more point needs to be emphasized about the relationship between electron configuration and the periodic table the atoms of elements in the same vertical column of the table have similar electron configurations for example consider the

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web the electron removed during the ionization of beryllium he 2s 2 is an s electron whereas the electron removed during the ionization of boron he 2s 2 2p 1 is a p electron this results in a lower first ionization energy for boron even though its nuclear charge is greater by one proton

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web predictions from reliable sources have been used for these elements grayed out electron numbers indicate subshells filled to their maximum bracketed noble gas symbols on the left represent inner configurations that are the same in each period written out these are he 2 helium 1s 2 ne 10 neon 1s 2 2s 2 2p 6

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web q1 what is the electron configuration of the element potassium answer the atomic number of potassium is 19 a neutral potassium element has 19 electrons its electron configuration is 1s 2 2s 2 2p 6 3s 2 3p 6 4s 1 or simply put as 4s 1 q2 what is the total number of electrons that can be filled in s p and d subshells

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web by building up from hydrogen this table can be used to determine the electron configuration for any atom on the periodic table we will now construct the ground state electron configuration and orbital diagram for a selection of atoms in the first and second periods of the periodic table

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web electron configurations are a simple way of writing down the locations of all of the electrons in an atom as we know the positively charged protons in the nucleus of an atom tend to attract negatively charged electrons

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