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# My Very First Of Motion

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 $\text{Cu} + 2\text{AgNO}_3 \rightarrow 2\text{Ag} + \text{Cu(NO}_3)_2$

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 $3\text{CO} + 2\text{H}_2\text{O} \rightarrow \text{C}_3\text{H}_8 + 2\text{H}_2\text{O}$

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web start with 8.3 mole  $\text{H}_2\text{S}$  and multiply it with 2 mole  $\text{H}_2\text{O}$  over 2 mole  $\text{H}_2\text{S}$  from the equation  $8.3 \text{ mole H}_2\text{S} \times 2 \text{ mole H}_2\text{O} / 2 \text{ mole H}_2\text{S} \times 18 \text{ g H}_2\text{O} = 149.4 \text{ g H}_2\text{O}$  is your theoretical yield so percent yield  $137.1 \text{ g H}_2\text{O} / 149.4 \text{ g H}_2\text{O} \times 100 = 91.77\%$  problems 6-8 are similar to problem 5

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web 1 csun.edu/hcchm001/introchemhandouts.html limiting reagents theoretical actual and percent yields example of a limiting reagent problem how many grams of  $\text{NH}_3$  can be produced theoretically from the reaction of 5.0 g of  $\text{N}$  what is the limiting reagent if 8.52 g are actually formed what is the

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web a write a balanced equation for the reaction  $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$  b determine the theoretical yield of  $\text{Al}_2\text{O}_3$  160.0 g  $\text{Al} \times 1 \text{ mol} / 2 \text{ mol} = 0.5 \text{ mol}$   $\text{Al}_2\text{O}_3 \times 101.964 \text{ g/mol} = 50.982 \text{ g}$   $\text{Al} \times 4 \text{ mol} / 1 \text{ mol} = 4 \text{ mol}$   $\text{Al}_2\text{O}_3 \times 101.964 \text{ g/mol} = 403.856 \text{ g}$  c determine the percent yield yield actual  $\times 100 = 260.0 \text{ g} / 403.856 \text{ g} \times 100 = 64.5\%$  theoretical 302.3 g

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web 7 na 2 c 2 o 4 is the limiting reactant percent yield 86 6 8 only four molecules can be made 9 this amount cannot be weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12 c is the limiting reagent 4 33 g of h 2 are left over 13

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web the percent yield is calculated as follows text percent yield  $\frac{\text{actual yield}}{\text{theoretical yield}} \times 100$  percent yield theoretical yieldactual yield 100 based on this definition we would expect a percent yield

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