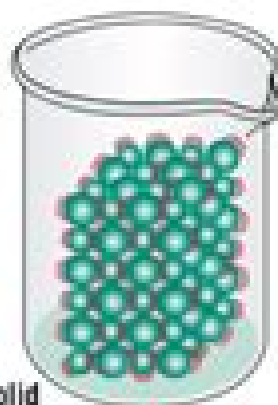


# The particle model

## Properties of solids:

- have a definite shape
- do not flow
- virtually impossible to compress
- expand if heated, but usually less than liquids and gases.

solid



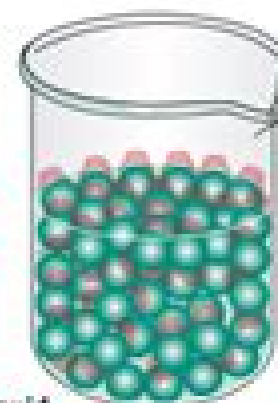
## Particles in solids:

- strongly bonded to each other
- vibrate a little, but not much compared to liquids and gases
- vibrate faster when heated.

## Properties of liquids:

- no definite shape
- can flow to take the shape of the bottom of a container
- very difficult to compress (virtually incompressible).

liquid



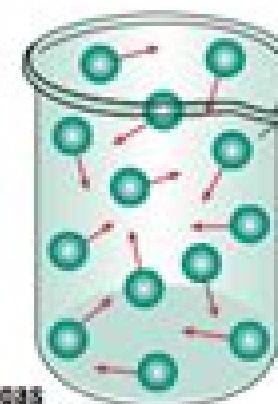
## Particles in liquids:

- weakly bonded to each other
- break their bonds easily
- vibrate and move more than those in a solid
- move faster when heated.

## Properties of gases:

- no fixed shape
- gases spread (or diffuse) to completely fill a container
- gases are easily compressed.

gas



## Gas particles:

- are 'free', having no bonds between them
- have much more energy than those of a solid or liquid
- fly around, bouncing off each other and the walls of their container.

# Particle Mechanics The Theory Of Energy States

**Jicheng Xie**



## **Particle Mechanics The Theory Of Energy States:**

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