

Reactions of Amines and Epoxydes

Preparation of Amines

Ammonium hydroxide: $\text{NH}_4^+ + \text{OH}^- \rightarrow \text{NH}_3 + \text{H}_2\text{O}$



Four different types of bonds



There are four types of bonding: ionic bonding and covalent bonding.

- Ionic bonds are formed by the transfer of electrons from one atom to another (anions and cations) (opposite charges).



Covalent ionic and polar covalent bonds

- Covalent bonds are classified as ionic (polar) and non-polar.
- Ionic bonds are formed by the transfer of electrons.
- Non-polar bonds are formed by the sharing of electrons.
- Polar covalent bonds are formed by the sharing of electrons.
- Non-polar covalent bonds are formed by the sharing of electrons.



Non-polar covalent bonds are stronger than ionic.

Polar and Nonpolar Covalent Bonds

Depending on the difference in the electronegativity values, polar bonds can be polar and nonpolar.

- If the difference in electronegativity is less than 0.4, the electrons are shared equally between the two atoms. Forming a nonpolar covalent bond (like a bond between chlorine and the C-C bond).
- If the difference in electronegativity is between 0.5 and 1.7, the more polar covalent bond (nonpolar covalent bond) electrons are shared in a region of N, O, F and other halogens.
- A difference of 1.7 or higher one or both of the atoms will have shared and unequal shared in a polar bond is formed.

A Summary of Ring-Opening Reactions of Epoxydes with Strong and Weak Protagonists

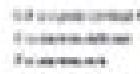


Protagonists and Reactions

Bond Types, Strength, Length, Polarity and Dipole Moment

Dipole Moment

The polarity of a bond is denoted by the dipole moment (μ).



Electronegativity is the ability to attract electrons to the atom.

- The dipole moment is shown as a vector pointing from the less electronegative atom to the more electronegative atom.
- The vector is proportional to the partial negative charge and the distance between a partial positive charge and a partial negative charge.
- The dipole plot indicates the charge difference.
- The dipole moment is shown as a distance and angle.

Molecular Dipole

To measure the molecular polarity, the sum of the dipole is used.

The polarity of the molecule is equal to the sum of the individual bond polarities. If dipole moments are canceled, a polar molecule is neutral.



Sum of the bond dipole moment for the molecule.

Bond Length and Bond Strength

Relative bond length



Length: C-C > C-O > C=O > H-O

Strength: C-C < C-O < C=O < H-O

Electronegativity: C < O < =O < H

Pauli exclusion principle: C < O < =O < H

London dispersion forces: C < O < =O < H

Hybridization: C < O < =O < H

Electronegativity: C < O < =O < H

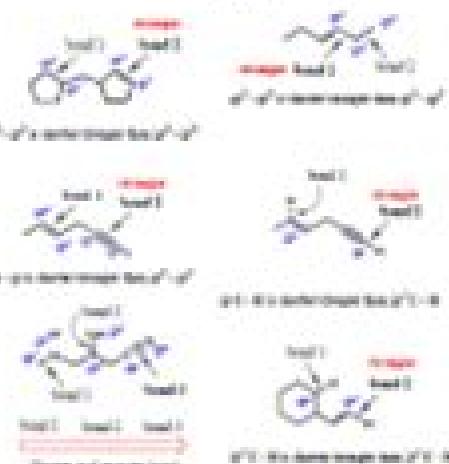
Relative bond strength

Average values of C-C and C-O bond depending on hybridization

| C-C Bond | $\text{H}_2\text{C}-\text{CH}_3$ | $\text{CH}_2=\text{CH}_2$ | CH_3CH_3 |
|---------------|----------------------------------|---------------------------|--------------------------|
| Hybridization | sp ³ hybrid | sp ² hybrid | sp ³ hybrid |
| Length | 1.17 Å | 1.35 Å | 1.35 Å |
| Strength | 1.12 kcal/mol | 1.11 kcal/mol | 1.07 kcal/mol |
| Hybridization | sp ³ | sp ² | sp ³ |
| Efficiency | 25.7% | 33.7% | 25.7% |
| Length | 1.10 Å | 1.19 Å | 1.10 Å |
| Strength | 0.91 kcal/mol | 0.91 kcal/mol | 0.91 kcal/mol |

Hybridization, dipole and strength bond

Example



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