

MOLE CONCEPT

Scientific notation
 $N \times 10^n$ → Integer
 $N < 10$

Density → $\frac{\text{mass}}{\text{volume}}$

Temperature →

$$^{\circ}\text{F} = \frac{9}{5} ({}^{\circ}\text{C}) + 32$$

$$^{\circ}\text{K} = {}^{\circ}\text{C} + 273.15$$

molar mass = mass of one mole substance in grams.

one mole: 6.022×10^{23} particles.

percentage composition: $\frac{\text{mass}_c}{\text{molar mass}_c} \times 100$

Empirical formula = molecular formula

limiting Reagent:

The reactant which present in less quantity.

gets consumed after sometime. Given mole by this we can find LR.

Isotopes - Z same A diff

Isobars - Z diff A same

Isoelectronic - Same num e⁻

Conversions

mass

$\frac{\text{mass}}{\text{molar mass}}$

6.022×10^{23}

$\div 6.022 \times 10^{23}$

Particals

Mole

$22.4 \times \frac{1}{\text{Molar mass}}$

Volume

at STP
 0°C 1 atm

Empirical formula: gives simplest whole number ratio of various elements in one molecule of compound.

Empirical formula =

formula of $\text{C}_2\text{H}_4\text{O}_6$
 empirical for. $\text{C}_2\text{H}_4\text{O}$

Molecular

formula =

$$MF = n \times EF$$

n is a common factor

gives actual ratio of atoms of various elements in one molecule of com.



H

O

* Mass per cent

$$= \frac{\text{mass of solute}}{\text{mass of solution}} \times 100$$

* Mole fraction

$$= \frac{\text{No. of moles of A}}{\text{No. of moles of solution}}$$

* Molarity

$$M = \frac{\text{No. of mole of solute}}{\text{volume of solution (lit)}}$$

* Molality

$$m = \frac{\text{no. of moles of solute}}{\text{mass of solvent (in kg)}}$$

Mole Notes

Robert Sydney Lowrie

Mole Notes:

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Unveiling the Energy of Verbal Art: An Mental Sojourn through **Mole Notes**

In a world inundated with monitors and the cacophony of instantaneous transmission, the profound energy and psychological resonance of verbal artistry often disappear into obscurity, eclipsed by the constant onslaught of sound and distractions.

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